

3: Definitions of Terms Relating to Individual Macromolecules, Their Assemblies, and Dilute Polymer Solutions (1988)

PREAMBLE

This document is part of a series published by the Commission on Macromolecular Nomenclature dealing with definitions for the important terms in polymer science [1-3]. It is also the first of several documents that will be published dealing specifically with physicochemical terminology in the polymer field. The recommendations presented here deal with such key areas of the physical chemistry of macromolecules as individual macromolecules, their assemblies and dilute polymer solutions; they include recommended terminology for molecular weight, molecular-weight averages, distribution functions, radius of gyration, the Flory-Huggins theory, viscosity of solutions, scattering of radiation by polymers, fractionation, etc.

The reader's attention is especially directed to the new terms 'uniform polymer' and 'nonuniform polymer' which denote polymers composed of molecules that are uniform or nonuniform, respectively, with respect to relative molecular mass and constitution. These terms replaced the widely used, but non-descriptive and self-contradictory terms 'monodisperse polymer' and 'polydisperse polymer'.

1 INDIVIDUAL MACROMOLECULES

<i>Term</i>	<i>Definition</i>
1.1 relative molecular mass, molecular weight	recommended symbol: M_r The ratio of the average mass per formula unit of a substance to 1/12 of the mass of an atom of nuclide ^{12}C .
	<i>Note</i> See Definition 1.2.
1.2 molar mass	recommended symbol: M Mass divided by amount of substance.
	<i>Notes</i> 1. Molar mass is usually expressed in g/mol or kg/mol units. The g/mol unit is recommended in polymer science, since then the numerical values of the molar mass and the relative molecular mass of a substance are equal.

Prepared by a working group consisting of P. Kratochvíl (Czechoslovakia) and U.W. Suter (Switzerland).
Reprinted from *Pure Appl. Chem.* **61**, 211-241 (1989).

DEFINITIONS

Term

Definition

2. Relative molecular mass (molecular weight) is a pure number and must not be associated with any units.
3. The use of the dalton as a unit of mass, identical with the atomic mass unit, is discouraged.
4. The terms 'molar' and 'molecular' may also be used for particles consisting of more than one molecule, such as complexes, aggregates, micelles, etc.
5. If there is no danger of confusion, the subscript *r* in the recommended symbol for the relative molecular mass, M_r , may be omitted.
- 1.3 degree of polymerization
recommended symbol: X
recommended abbreviation: DP
The number of monomeric units in a molecule of a polymer.
- 1.4 thermodynamically equivalent sphere
A sphere, impenetrable to other spheres, displaying the same excluded volume (see Definition 3.1.9) as an actual polymer molecule.
- 1.5 short-range intramolecular interaction
Steric or other interaction involving atoms or groups or both situated nearby in sequence along the chain.

Notes

1. The interacting atoms or groups are typically separated by fewer than ten consecutive bonds in a chain.
2. If no confusion can occur, the word 'intramolecular' may be omitted.
- 1.6 long-range intramolecular interaction
Interaction between segments, widely separated in sequence along the chain, that occasionally approach one another during molecular flexing.

Notes

1. This type of interaction is closely related to the excluded volume of a segment (see Definition 3.1.8), the latter quantity reflecting interactions involving segments and solvent molecules.
2. If no confusion can occur, the word 'intramolecular' may be omitted.
- 1.7 unperturbed dimensions
The dimensions of an actual polymer random coil (see Definition 1.15) in a theta state (see Definition 3.1.5).
- 1.8 perturbed dimensions
The dimensions of an actual polymer random coil (see Definition 1.15) not in a theta state (see Definition 3.1.5).
- 1.9 radius of gyration
recommended symbols: s , $\langle s^2 \rangle^{1/2}$
A parameter characterizing the size of a particle of any shape. For a rigid particle consisting of mass elements of mass m_i ,

Term

Definition

each located at a distance r_i from the centre of mass, the radius of gyration, s , is defined as the square root of the mass-average of r_i^2 for all the mass elements, i.e.

$$s = \left(\frac{\sum_i m_i r_i^2}{\sum_i m_i} \right)^{1/2}$$

For a non-rigid particle, an average over all conformations is considered, i.e.

$$\langle s^2 \rangle^{1/2} = \left\langle \left(\frac{\sum_i m_i r_i^2}{\sum_i m_i} \right)^{1/2} \right\rangle$$

Note

The subscript zero is used to indicate unperturbed dimensions (see Definition 1.7), as in $\langle s^2 \rangle_0^{1/2}$.

1.10 end-to-end vector

recommended symbol: r

The vector connecting the two ends of a linear polymer chain in a particular conformation.

1.11 end-to-end distance

recommended symbol: r

The length of the end-to-end vector (see Definition 1.10).

1.12 root-mean-square end-to-end distance

recommended symbol: $\langle r^2 \rangle^{1/2}$

The square root of the mean-square end-to-end distance (see Definition 1.11) of a linear polymer chain averaged over all conformations of the chain. For a freely jointed chain (see Definition 1.16) consisting of N segments each of length L , $\langle r^2 \rangle^{1/2} = N^{1/2} L$.

Notes

1. The subscript zero is used to indicate unperturbed dimensions (see Definition 1.7), as in $\langle r^2 \rangle_0^{1/2}$
2. If this term is used repeatedly, and if it is not confusing, the abbreviated name 'end-to-end distance' may be used.

1.13 characteristic ratio

recommended symbol: C_N (C_∞ when $N \rightarrow \infty$)

The ratio of the mean-square end-to-end distance (see Definition 1.12), $\langle r^2 \rangle_0$ of a linear polymer chain in a theta state (see Definition 3.1.5) to $N \cdot L^2$, where N is the number of rigid sections in the main chain, each of length L ; if all of the rigid sections are not of equal length, the mean-square value of L is used, i.e.

$$L^2 = \sum_i \bar{L}_i^2 / N$$

Note

In simple single-strand chains, the bonds are taken as the rigid sections.

DEFINITIONS

<i>Term</i>	<i>Definition</i>
1.14 contour length	<p>The maximum end-to-end distance (see Definition 1.11) of a linear polymer chain.</p> <p><i>Notes</i></p> <ol style="list-style-type: none">1. For a single-strand polymer molecule, this usually means the end-to-end distance of the chain extended to the all-trans conformation. For chains with complex structure, only an approximate value of the contour length may be accessible.2. The sum of the lengths of all skeletal bonds of a single-strand polymer molecule is occasionally termed 'contour length'. This use of the term is discouraged.
1.15 random coil, statistical coil	<p>The complete set of spatial arrangements of a chain molecule with a large number of segments that randomly change mutual orientation with time, under conditions in which it is free from external constraints that would affect its conformation.</p> <p><i>Note</i></p> <p>If the solution of the chain molecules is not in a theta state (see Definition 3.1.5), the segments change mutual orientation only approximately randomly.</p>
1.16 freely jointed chain	<p>A hypothetical linear chain molecule consisting of infinitely thin rectilinear segments uniform in length; each segment can take all orientations in space with equal probability, independently of its neighbours.</p> <p><i>Note</i></p> <p>For models in which the segments are not all uniform in length, the name 'random-walk chain' has been used.</p>
1.17 equivalent chain	<p>A hypothetical freely jointed chain (see Definition 1.16) with the same mean-square end-to-end distance (see Definition 1.11) and contour length (see Definition 1.14) as an actual polymer chain in a theta state (see Definition 3.1.5).</p>
1.18 statistical segment	<p>A segment of an actual polymer chain which behaves, with respect to some property, virtually as a segment of a freely jointed chain (see Definition 1.16).</p>
1.19 freely rotating chain	<p>A hypothetical linear chain molecule, free from short-range and long-range interactions (see Definitions 1.5 and 1.6), consisting of infinitely thin rectilinear segments (bonds) of fixed length, jointed at fixed bond angles; the torsion angles of the bonds can assume all values with equal probability.</p>
1.20 steric factor	<p>recommended symbol: σ</p> <p>The ratio of the root-mean-square end-to-end distance (see</p>

<i>Term</i>	<p><i>Definition</i></p> <p>Definition 1.12) of a polymer chain with unperturbed dimensions (see Definition 1.7), $\langle r^2 \rangle_0^{1/2}$, to that of a freely rotating chain (see Definition 1.19) with the same structure, $\langle r^2 \rangle_{0,f}^{1/2}$, i.e., $(\langle r^2 \rangle_0 / \langle r^2 \rangle_{0,f})^{1/2}$, in the limit of infinite chain length.</p> <p><i>Note</i></p> <p>The steric factor reflects the effect of hindrance to free rotation.</p>
1.21 worm-like chain, continuously curved chain	<p>A hypothetical linear macromolecule consisting of an infinitely thin chain of continuous curvature; the direction of curvature at any point is random.</p> <p><i>Notes</i></p> <ol style="list-style-type: none"> 1. The model describes the whole spectrum of chains with different degrees of chain stiffness from rigid rods to random coils (see Definition 1.15), and is particularly useful for representing stiff chains. 2. In the literature this chain is sometimes referred to as Porod-Kratky chain.
1.22 persistence length	<p>recommended symbol: a</p> <p>The average projection of the end-to-end vector (see Definition 1.10) on the tangent to the chain contour at a chain end in the limit of infinite chain length.</p> <p><i>Note</i></p> <p>The persistence length is the basic characteristic of the worm-like chain (see Definition 1.21).</p>
1.23 short-chain branch	An oligomeric offshoot from a macromolecular chain.
1.24 long-chain branch	A polymeric offshoot from a macromolecular chain.
1.25 branching index	<p>recommended symbol: g</p> <p>A parameter characterizing the effect of long-chain branches (see Definition 1.24) on the size of a branched macromolecule in solution and defined as the ratio of the mean-square radius of gyration (see Definition 1.9) of a branched molecule, $\langle s_b^2 \rangle$, to that of an otherwise identical linear molecule, $\langle s_l^2 \rangle$, with the same relative molecular mass (see Definition 1.1) in the same solvent and at the same temperature, i.e. $g = \langle s_b^2 \rangle / \langle s_l^2 \rangle$.</p>
1.26 network	A highly ramified polymer structure in which each constitutional unit is connected to each other constitutional unit and to the macroscopic phase boundary by many permanent paths through the structure, their number increasing with the average number

DEFINITIONS

Term *Definition*
of intervening bonds; these paths must on the average be coextensive with this structure.

Note

Usually, and in all systems that exhibit rubber elasticity, the number of distinct paths is very high, but some constitutional units exist, in most cases, which are connected by a single path only. Sometimes, a structure without any multiple path has also been called a network.

1.27 microgel A network of microscopic dimensions.

1.28 copolymer micelle A micelle formed by one or more block or graft copolymer molecules in a selective solvent (see Definition 3.1.14).

2 ASSEMBLIES OF MACROMOLECULES

Term *Definition*
2.1 compositional heterogeneity The variation in elemental composition from molecule to molecule usually found in copolymers.

2.2 constitutional heterogeneity The variation in constitution from molecule to molecule in polymers with molecules uniform with respect to elemental composition.

Note

An example is a polymer composed of linear and branched molecules; another example is a statistical copolymer comprising two isomeric constitutional units.

2.3 uniform polymer, monodisperse polymer A polymer composed of molecules uniform with respect to relative molecular mass (see Definition 1.1) and constitution.

Notes

1. A polymer comprising a mixture of linear and branched chains, all of uniform relative molecular mass, is not uniform.
2. A copolymer comprising linear molecules of uniform relative molecular mass and uniform elemental composition, but different sequence arrangement of the various types of monomeric units, is not uniform (e.g. a copolymer comprising molecules with random arrangement as well as block arrangement of monomeric units).
3. A polymer uniform with respect only to either relative molecular mass or constitution may be termed 'uniform', provided a suitable qualifier is used (e.g. 'a polymer uniform with respect to relative molecular mass').
4. The adjectives 'monodisperse' and 'polydisperse' (see Definition

<i>Term</i>	<i>Definition</i>
	<p>2.4) are deeply rooted in the literature despite being non-descriptive and self-contradictory. They are in common usage and it is recognized that they will continue to be used for a certain time, nevertheless more satisfactory terms are clearly desirable. After an extensive search for possible replacements, the new terms 'uniform' and 'non-uniform'(see Definition 2.4) have been selected and they are now the preferred adjectives.</p>
2.4 non-uniform polymer, polydisperse polymer	<p>A polymer comprising molecules non-uniform with respect to relative molecular mass (see Definition 1.1) or constitution or both.</p> <p><i>Note</i> See Definition 2.3, <i>Note</i> 3.</p>
2.5 molar-mass average, relative molecular mass average/molecular-weight average	<p>recommended symbols: \bar{M}_k and $\bar{M}_{r,k}$, respectively In both symbols, k specifies the type of average. Any average of the molar mass (see Definition 1.2) or relative molecular mass (molecular weight) (see Definition 1.1) for a non-uniform polymer.</p> <p><i>Notes</i> 1. An infinite number of molar-mass averages can in principle be defined, but only a few types of averages are directly accessible experimentally. The most important averages are defined by simple moments of the distribution functions (see Definition 2.13) and are obtained by methods applied to systems in thermodynamic equilibrium, such as osmometry, light scattering and sedimentation equilibrium. Hydrodynamic methods, as a rule, yield more complex molar-mass averages. 2. Any molar-mass average can be defined in terms of mass fractions or mole fractions. In this document only a few of the important molar-mass averages are given in terms of the mass fractions, w_i, of the species with molar mass M_i. These definitions are most closely related to the experimental determination of molar-mass averages.</p>
2.6 number-average molar mass	<p>recommended symbol: \bar{M}_n</p> $\bar{M}_n = \frac{1}{\sum_i (w_i/M_i)}$
number-average relative molecular mass, number-average molecular weight	<p>recommended symbol: $\bar{M}_{r,n}$</p> $\bar{M}_{r,n} = \frac{1}{\sum_i (w_i/M_{r,i})}$ <p>For explanation of symbols, see Definition 2.5.</p>

DEFINITIONS

<i>Term</i>	<i>Definition</i>
2.7 mass-average molar mass	recommended symbol: \bar{M}_w $\bar{M}_w = \sum_i w_i M_i$
mass-average relative molecular mass, weight-average molecular weight	recommended symbol: $\bar{M}_{r,w}$ $\bar{M}_{r,w} = \sum_i w_i M_{r,i}$ For explanation of symbols, see Definition 2.5.
2.8 z-average molar mass	recommended symbol: \bar{M}_z $\bar{M}_z = \frac{\sum_i w_i M_i^2}{\sum_i w_i M_i}$
z-average relative molecular mass, z-average molecular weight	recommended symbol: $\bar{M}_{r,z}$ $\bar{M}_{r,z} = \frac{\sum_i w_i M_{r,i}^2}{\sum_i w_i M_{r,i}}$ For explanation of symbols, see Definition 2.5.
2.9 (z+1)-average molar mass	recommended symbol: \bar{M}_{z+1} $\bar{M}_{z+1} = \frac{\sum_i w_i M_i^3}{\sum_i w_i M_i^2}$
(z+1)-average relative molecular mass, (z+1)-average molecular weight	recommended symbol: $\bar{M}_{r,z+1}$ $\bar{M}_{r,z+1} = \frac{\sum_i w_i M_{r,i}^3}{\sum_i w_i M_{r,i}^2}$
2.10 viscosity-average molar mass	For explanation of symbols, see Definition 2.5. recommended symbol: \bar{M}_v $\bar{M}_v = \left[\sum_i w_i M_i^a \right]^{1/a}$
viscosity-average relative molecular mass, viscosity-average molecular weight	recommended symbol: $\bar{M}_{r,v}$ $\bar{M}_{r,v} = \left[\sum_i w_i M_{r,i}^a \right]^{1/a}$ where a is the exponent in the Mark–Houwink equation (see Definition 3.2.25), $[\eta] = K \cdot M^a$; for explanation of the other symbols, see Definition 2.5.

<i>Term</i>	<i>Definition</i>
	<i>Note</i> The exponent a is not identical with the adjustable parameter of some of the distribution functions (see Definitions 2.16-2.20) or with the persistence length (see Definition 1.22).
2.11 apparent molar mass	recommended symbol: M_{app}
apparent relative molecular mass, apparent molecular weight	recommended symbol: $M_{\text{r,app}}$ Molar mass, relative molecular mass, or molecular weight calculated from experimental data without the application of appropriate corrections, such as for finite polymer concentration, association, preferential solvation, compositional heterogeneity (see Definition 2.1), constitutional heterogeneity (see Definition 2.2).
2.12 average degree of polymerization	recommended symbol: \bar{X}_k where k specifies the type of average. Any average of the degree of polymerization (see Definition 1.3) for a polymer. <i>Note</i> Definitions 2.5-2.10 apply directly to averages of the degree of polymerization when X is substituted for M in the formulae.
2.13 distribution function	A normalized function giving the relative amount of a portion of a polymeric substance with a specific value, or a range of values, of a random variable or variables. <i>Notes</i> 1. Distribution functions may be discrete, i.e. take on only certain specified values of the random variable(s), or continuous, i.e. take on any intermediate value of the random variable(s), in a given range. Most distributions in polymer science are intrinsically discrete, but it is often convenient to regard them as continuous or to use distribution functions that are inherently continuous. 2. Distribution functions may be integral (or cumulative), i.e. give the proportion of the population for which a random variable is less than or equal to a given value. Alternatively they may be differential distribution functions (or probability density functions), i.e. give the (maybe infinitesimal) proportion of the population for which the random variable(s) is (are) within a (maybe infinitesimal) interval of its (their) range(s). 3. Normalization requires that: (i) for a discrete differential distribution function, the sum of the function values over all possible values of the random variable(s) be unity; (ii) for a continuous differential distribution function, the integral over the entire range of the random variable(s) be unity; (iii) for an integral (cumulative) distribution function, the function value at the upper limit of the random variable(s) be unity.

DEFINITIONS

<i>Term</i>	<i>Definition</i>
2.14 number-distribution function	A distribution function (see Definition 2.13) in which the relative amount of a portion of a substance with a specific value, or a range of values, of the random variable(s) is expressed in terms of mole fraction.
2.15 mass-distribution function, weight-distribution function	A distribution function (see Definition 2.13) in which the relative amount of a portion of a substance with a specific value, or a range of values, of the random variable(s) is expressed in terms of mass fraction.
2.16 Schulz-Zimm distribution	A continuous distribution with the differential mass-distribution function (see Definition 2.15) of the form $f_w(w) dx = \frac{a^{b+1}}{\Gamma(b+1)} x^b \exp(-ax) dx,$ where x is a parameter characterizing the chain length, such as relative molecular mass (see Definition 1.1) or degree of polymerization (see Definition 1.3), a and b are positive adjustable parameters, and $\Gamma(b+1)$ is the gamma function of $(b+1)$.
2.17 most probable distribution	A discrete distribution with the differential mass-distribution function (see Definition 2.15) of the form $f_w(x) = a^2 x(1-a)^{x-1}.$ For explanation of symbols, see Definition 2.16. <p><i>Notes</i></p> <ol style="list-style-type: none"> 1. For large values of x, the most probable distribution converges to the particular case of the Schulz-Zimm distribution (see Definition 2.16) with $b=1$. 2. In the literature, this distribution is sometimes referred to as the Flory distribution or the Schulz-Flory distribution.
2.18 Poisson distribution	A discrete distribution with the differential mass-distribution function (see Definition 2.15) of the form $f_w(x) = \frac{x}{a+1} \frac{e^{-a} a^{x-1}}{(x-1)!}$ For explanation of symbols, see Definition 2.16.
2.19 Tung distribution	A continuous distribution with the differential mass-distribution function (see Definition 2.15) of the form $f_w(x) dx = abx^{b-1} \exp(ax^b) dx.$ For explanation of symbols, see Definition 2.16.

DEFINITIONS

<i>Term</i>	<i>Definition</i>
	<i>Notes</i> <ol style="list-style-type: none">1. In some respects, a polymer solution in the theta state resembles an ideal solution and the theta state may be referred to as a pseudo-ideal state. However, a solution in the theta state must not be identified with an ideal solution.2. In the literature, the theta state is also written 'θ state'.3. The solvent involved is often referred to as 'θ solvent'.4. It is assumed that the molar mass of the polymer is high.
3.1.6 theta temperature	The temperature at which a solution is in the theta state (see Definition 3.1.5).
	<i>Note</i> <p>In the literature, the theta temperature is also written 'θ temperature'.</p>
3.1.7 virial coefficients, virial coefficients of the chemical potential	recommended symbol: A_i , where $i = 1, 2$, etc. The coefficients in the expansion of the chemical potential of the solvent, \boldsymbol{m} , in powers of the mass concentration, c , of the solute, i.e. $\mu_s - \mu_s^0 = -\pi V_s = -RTV_s(A_1c + A_2c^2 + A_3c^3 + \dots),$ where \boldsymbol{m}^0 is the chemical potential of the solvent in the reference state at the temperature of the system and ambient pressure, π is the osmotic pressure and V_s is the partial molar volume of the solvent. In solvents comprising more than one component, the definition applies to any solvent component. The first virial coefficient is the reciprocal number-average molar mass (see Definition 2.6), i.e. $A_1 = 1/\bar{M}_n$. The second and higher virial coefficients, A_2, A_3, \dots , respectively, describe polymer-solvent (see Definition 3.1.3) and polymer-polymer interactions.
	<i>Note</i> <p>The factor RT is sometimes included in the virial coefficients.</p>
3.1.8 excluded volume of a segment	The volume from which a segment of a macromolecule in solution effectively excludes all other segments, i.e. those belonging to the same macromolecule as well as those belonging to other macromolecules.
	<i>Note</i> <p>The excluded volume of a segment depends on the Gibbs and Helmholtz energies of mixing of solvent and polymer, i.e. on the thermodynamic quality of the solvent (see Definition 3.1.4), and is not a measure of the geometrical volume of that segment.</p>
3.1.9 excluded volume of a macromolecule	The volume from which a macromolecule in a dilute solution effectively excludes all other macromolecules.

<i>Term</i>	<i>Definition</i>
	<i>Note</i> The excluded volume of a macromolecule depends on the Gibbs and Helmholtz energies of mixing of solvent and polymer, i.e. on the thermodynamic quality of the solvent (see Definition 3.1.4), and is not a measure of the geometrical volume of that macromolecule.
3.1.10 expansion factor	recommended symbols: α_r , α_s , α_η The ratio of a dimensional characteristic of a macromolecule in a given solvent at a given temperature to the same dimensional characteristic in the theta state (see Definition 3.1.5) at the same temperature. The most frequently used expansion factors are: expansion factor of the mean-square end-to-end distance (see Definition 1.12), $\alpha_r = (\langle r^2 \rangle / \langle r^2 \rangle_0)^{1/2}$; expansion factor of the radius of gyration (see Definition 1.9), $\alpha_s = (\langle s^2 \rangle / \langle s^2 \rangle_0)^{1/2}$; viscosity expansion factor $\alpha_\eta = ([\eta] / [\eta]_0)^{1/3}$, where $[\eta]$ and $[\eta]_0$ are the intrinsic viscosity (see Definition 3.2.21) in a given solvent and in the theta state at the same temperature, respectively.
	<i>Note</i> Expansion factors defined by different dimensional characteristics are not exactly equal, nor need they have a constant ratio as a function of relative molecular mass.
3.1.11 Flory-Huggins theory	A thermodynamic theory of polymer solutions, first formulated independently by Flory and by Huggins, in which the thermodynamic quantities of the solution are derived from a simple concept of combinatorial entropy of mixing and a reduced Gibbs-energy parameter, the 'c parameter' (see Definition 3.1.12).
3.1.12 c parameter	recommended symbol: c A numerical parameter employed in the Flory-Huggins theory (see Definition 3.1.11), which accounts in the main for the contribution of the non-combinatorial entropy of mixing and for the enthalpy of mixing.
3.1.13 preferential sorption, selective sorption	An equilibrium phenomenon, operative in polymer solutions in multicomponent solvents and in polymer networks (see Definition 1.26) swollen by multicomponent solvents, that produces differences in solvent composition in the polymer containing region and in the pure solvent which is in thermodynamic equilibrium with that region.
3.1.14 selective solvent	A medium that is a solvent for at least one component of a mixture of polymers, or for at least one block of a block or graft polymer, but a non-solvent for the other component(s) or block(s).

DEFINITIONS

<i>Term</i>	<i>Definition</i>
3.1.15 co-solvency	The dissolution of a polymer in a solvent comprising more than one component, each component of which by itself is a non solvent for the polymer.
3.1.16 solubility parameter (of a polymer)	recommended symbol: <i>d</i> A characteristic of a polymer used in predicting solubility of that polymer in a given solvent. <i>Notes</i> 1. For a substance of low molecular weight, the value of the solubility parameter is often estimated from the enthalpy of vaporization; for a polymer, it is usually taken to be the value of the solubility parameter of the solvent producing the solution with maximum intrinsic viscosity (see Definition 3.2.21) or maximum swelling of a network (see Definition 1.26) of the polymer. 2. The solubility parameter is usually expressed in $(\text{cal}/\text{cm}^3)^{1/2}$ or, preferably, $(\text{J}/\text{cm}^3)^{1/2}$ units. The units must always be given; $1(\text{cal}/\text{cm}^3)^{1/2} = 2.05(\text{J}/\text{cm}^3)^{1/2}$.
3.1.17 isopycnic	An adjective describing components of a multicomponent system with equal partial specific volumes.

3.2 Transport properties

<i>Term</i>	<i>Definition</i>
3.2.1 frictional coefficient	recommended symbol: <i>f</i> A tensor correlating the frictional force <i>F</i> , opposing the motion of a particle in a viscous fluid and the velocity <i>a</i> of this particle relative to the fluid. <i>Note</i> In the case of an isolated spherical particle in a viscous isotropic fluid, <i>f</i> is a constant and $F = fu$.
3.2.2 hydrodynamically equivalent sphere	A hypothetical sphere, impenetrable to the surrounding medium, displaying in a hydrodynamic field the same frictional effect as an actual polymer molecule. <i>Note</i> The size of a hydrodynamically equivalent sphere may be different for different types of motion of the macromolecule, e.g. for diffusion and for viscous flow.
3.2.3 hydrodynamic volume	The volume of a hydrodynamically equivalent sphere (see Definition 3.2.2).
3.2.4 bead-rod model	A model simulating the hydrodynamic properties of a chain

<i>Term</i>	<i>Definition</i>
	macromolecule consisting of a sequence of beads, each of which offers hydrodynamic resistance to the flow of the surrounding medium and is connected to the next bead by a rigid rod which does not. The mutual orientation of the rods is random.
3.2.5 bead-spring model	A model simulating the hydrodynamic properties of a chain macromolecule consisting of a sequence of beads, each of which offers hydrodynamic resistance to the flow of the surrounding medium and is connected to the next bead by a spring which does not contribute to the frictional interaction but which is responsible for the elastic and deformational properties of the chain. The mutual orientation of the springs is random.
3.2.6 freely draining	An adjective referring to a chain macromolecule the segments of which produce such small frictional effects when moving in a medium that the hydrodynamic field in the vicinity of a given segment is not affected by the presence of other segments. Thus, the solvent can flow virtually undisturbed through the domain occupied by a freely draining macromolecule.
3.2.7 non-draining	An adjective describing a chain macromolecule that behaves in a hydrodynamic field as though the solvent within the domain of the macromolecule were virtually immobilized with respect to the macromolecule.
3.2.8 partially draining	An adjective describing a chain macromolecule that behaves in a hydrodynamic field as though the solvent, within the domain of the macromolecule were progressively more immobilized with respect to the macromolecule in the direction from its outer fringes inward.
	<i>Note</i> A freely draining (see Definition 3.2.6) macromolecule and a non-draining (see Definition 3.2.7) macromolecule are two extremes of the concept of a partially draining macromolecule.
3.2.9 streaming birefringence, flow birefringence	The birefringence induced by flow in liquids, solutions and dispersions of optically anisotropic, anisometric or deformable molecules or particles due to a non-random orientation of the molecules or particles.
3.2.10 rotational diffusion	A process by which the equilibrium statistical distribution of the overall orientation of molecules or particles is maintained or restored.

Note
Rotational diffusion may be compared to translational diffusion,

DEFINITIONS

<i>Term</i>	<i>Definition</i>
3.2.11 sedimentation coefficient	<p>through which the equilibrium statistical distribution of position in space is maintained or restored.</p> <p>recommended symbol: s A parameter characterizing the motion of a particle in a centrifugal field and defined as the velocity of motion u due to unit centrifugal acceleration, i.e. $s = u/(r \cdot \omega^2)$, where ω is the angular velocity and r the distance from the centre of rotation.</p> <p><i>Note</i> The unit 10^{-13} second is useful; this unit has been referred to as a 'svedberg' which, however, is not an SI unit.</p>
3.2.12 sedimentation equilibrium	<p>The equilibrium established in a centrifugal field when there is no net flux of any component across any plane perpendicular to the centrifugal force.</p>
3.2.13 equilibrium sedimentation (method)	<p>A method by which the distribution of the concentration of the solute or dispersed component in a dilute solution or dispersion along the centrifuge cell is measured at sedimentation equilibrium (see Definition 3.2.12), and the results are interpreted in terms of molar masses (see Definition 1.2) or their distribution, or both.</p>
3.2.14 sedimentation velocity method	<p>A method by which the velocity of motion of solute component(s) or dispersed particles is measured and the result is expressed in terms of its (their) sedimentation coefficient(s) (see Definition 3.2.11).</p>
3.2.15 Archibald's method	<p>A sedimentation method based on the fact that at the meniscus and at the bottom of the centrifuge cell there is never a flux of the solute across a plane perpendicular to the radial direction and the equations characterizing the sedimentation equilibrium (see Definition 3.2.12) always apply there, even though the system as a whole may be far from equilibrium.</p> <p><i>Note</i> The use of the term 'approach to sedimentation equilibrium' for Archibald's method is discouraged, since it has a more general meaning.</p>
3.2.16 equilibrium sedimentation in a density gradient	<p>An equilibrium sedimentation technique (see Definition 3.2.13) working with a multi-component solvent forming a density gradient in a centrifugal field.</p>
3.2.17 relative viscosity, viscosity ratio	<p>recommended symbol: η_r The ratio of the viscosity of the solution, η, to the viscosity of the solvent, η_s, i.e. $\eta_r = \eta/\eta_s$.</p>

<i>Term</i>	<i>Definition</i>
3.2.18 relative viscosity increment	<p>recommended symbol: η_i</p> <p>The ratio of the difference between the viscosities of solution and solvent to the viscosity of the solvent, i.e. $\eta_i = (\eta - \eta_s)/\eta_s$. For explanation of symbols, see Definition 3.2.17.</p> <p><i>Note</i></p> <p>The use of the term 'specific viscosity' for this quantity is discouraged, since the relative viscosity increment does not have the attributes of a specific quantity.</p>
3.2.19 reduced viscosity, viscosity number	<p>The ratio of the relative viscosity increment (see Definition 3.2.18) to the mass concentration of the polymer, c, i.e. η_i/c.</p> <p><i>Notes</i></p> <ol style="list-style-type: none"> 1. The unit must be specified; cm^3/g is recommended. 2. This quantity and those in Definitions 3.2.20 and 3.2.21 are neither viscosities nor pure numbers. The terms are to be looked on as traditional names. Any replacement by consistent terminology would produce unnecessary confusion in the polymer literature.
3.2.20 inherent viscosity, logarithmic viscosity number	<p>recommended symbols: η_{inh}, η_{ln}, respectively</p> <p>The ratio of the natural logarithm of the relative viscosity (see Definition 3.2.17) to the mass concentration of the polymer, c, i.e.</p> $\eta_{\text{inh}} \equiv \eta_{\text{ln}} = (\ln \eta_r)/c.$ <p><i>Notes</i></p> <p>See notes under Definition 3.2.19.</p>
3.2.21 intrinsic viscosity, limiting viscosity number	<p>recommended symbol: $[\eta]$</p> <p>The limiting value of the reduced viscosity (see Definition 3.2.19) or the inherent viscosity (see Definition 3.2.20) at infinite dilution of the polymer, i.e.</p> $[\eta] = \lim_{c \rightarrow 0} (\eta_i/c) = \lim_{c \rightarrow 0} \eta_{\text{inh}}$ <p><i>Notes</i></p> <ol style="list-style-type: none"> 1. This term is also known in the literature as the Staudinger index. 2. See notes under Definition 3.2.19.
3.2.22 Huggins equation	<p>The equation describing the dependence of the reduced viscosity (see Definition 3.2.19), η_i/c, on the mass concentration of the polymer, c, for dilute polymer solutions (see Definition 3.1.1) of the form</p> $\eta_i/c = [\eta] + k_{\text{H}}[\eta]^2 c$

DEFINITIONS

<i>Term</i>	<i>Definition</i>
	where k_H is the Huggins coefficient (see Definition 3.2.23) and $[\eta]$ is the intrinsic viscosity (see Definition 3.2.21).
3.2.23 Huggins coefficient	recommended symbol: k_H A parameter in the Huggins equation (see Definition 3.2.22).
3.2.24 viscosity function	recommended symbol: Φ A coefficient connecting the intrinsic viscosity (see Definition 3.2.21), the radius of gyration (see Definition 1.9) and the molar mass (see Definition 1.2) of a chain macromolecule, according to the equation $[\eta] = \Phi \overline{r_g^2}^{3/2} / M$.
	<i>Note</i> The viscosity function is often referred to as the Flory constant.
3.2.25 Mark-Houwink equation	The equation describing the dependence of the intrinsic viscosity (see Definition 3.2.21) of a polymer on its relative molecular mass (molecular weight) (see Definition 1.1) and having the form $[\eta] = K \cdot M_r^a$ where K and a are constants, the values of which depend on the nature of the polymer and solvent as well as on temperature; M_r , is usually one of the relative molecular-mass averages (see Definition 2.5). <i>Notes</i> <ol style="list-style-type: none">1. The use of this equation with the relative molecular mass (molecular weight) is recommended, rather than with molar mass (which has the dimension of mass divided by amount of substance), since in the latter case the constant K assumes awkward and variable dimensions owing to the fractional and variable nature of the exponent a.2. Kuhn and Sakurada have also made important contributions and their names are sometimes included, as, for example, in the Kuhn-Mark-Houwink-Sakurada equation.

3.3 Coherent elastic scattering of radiation

A beam of radiation traversing a medium may be attenuated and partially scattered. The definitions below are for those cases in which the attenuation of the incident beam is due only to scattering, the energy of scattering quanta is the same as that of quanta in the primary beam (elastic scattering) and phase relationships between independent scatterers are retained (coherent scattering). This document deals with light scattering (LS), small-angle x-ray scattering (SAXS), and small-angle neutron scattering (SANS). In light scattering the polarization of light is relevant; plane-polarized light is considered here only, and it is called vertically polarized (v) if the electric vector of the beam is perpendicular to the plane containing the source, sample and detector, and horizontally polarized (h) if the electric vector lies in that plane. Unpolarized light is considered to be a mixture of equal parts of v and h light.

INDIVIDUAL MACROMOLECULES

<i>Term</i>	<i>Definition</i>
3.3.1 small particle	A particle much smaller than the wavelength of the radiation in the medium. In practice, all dimensions of a particle considered small must be less than about one-twentieth of the wavelength employed.
3.3.2 large particle	A particle with dimensions comparable with the wavelength of the radiation in the medium or larger. In practice a particle must be treated as large, if its largest dimension exceeds about one-twentieth of the wavelength employed.
3.3.3 scattering angle, angle of observation	recommended symbol: \mathbf{q} The angle between the forward direction of the incident beam and a straight line connecting the scattering point and the detector.
3.3.4 scattering vector	Vector difference between the wave propagation vectors of the incident and the scattered beam, both of length $2\delta/\bar{\epsilon}$, where $\bar{\epsilon}$ is the wavelength of the scattered radiation in the medium.
3.3.5 length of the scattering vector	recommended symbol: q The length of the scattering vector (see Definition 3.3.4) is $q = (4\pi/\lambda) \cdot \sin(\theta/2)$, where λ is the wavelength of the scattered radiation in the medium and θ is the scattering angle (see Definition 3.3.3).
3.3.6 refractive index increment	recommended symbol: $\partial n/\partial C$ The change of the solution refractive index, n , with solute concentration, C .
	<i>Notes</i> 1. The solute concentration is most frequently expressed in terms of mass concentration, molality or volume fraction. If expressed in terms of mass concentration or molality, the corresponding refractive index increments are referred to as specific or molal refractive index increments, respectively. 2. Following use of the full name, the abbreviated name refractive increment may be used.
3.3.7 Rayleigh ratio	recommended symbols: $R(\theta)$, R_θ The quantity used to characterize the scattered intensity at the scattering angle θ (see Definition 3.3.3), defined as $R(\theta) = i_\theta r^2 / (I \cdot f \cdot V)$, where I is the intensity of the incident radiation, i_θ is the total intensity of scattered radiation observed at an angle θ and a distance r from the point of scattering and V is the scattering volume. The factor f takes account of polarization phenomena.

DEFINITIONS

Term

Definition

It depends on the type of radiation employed:

1. for light scattering, dependent on the polarization of the incident beam, $f = 1$ for vertically polarized light, $f = \cos^2 \theta$ for horizontally polarized light, $f = (1 + \cos^2 \theta)/2$ for unpolarized light;
2. for small-angle neutron scattering $f = 1$;
3. for small-angle x-ray scattering $f \approx 1$, if $\theta < \text{ca. } 5^\circ$.

Notes

1. The dimension of $R(\theta)$ is $(\text{length})^{-1}$ and the unit commonly used is cm^{-1} .
2. In small-angle neutron scattering the term cross-section is often used instead of $R(\theta)$; the two quantities are identical.

3.3.8 excess Rayleigh ratio

The difference between the Rayleigh ratio (see Definition 3.3.7) for a dilute solution and for pure solvent.

Note

If the scattering intensity is not reduced to the Rayleigh ratio, the difference between the scattering intensities for a dilute solution and that for pure solvent is named 'excess scattering'.

3.3.9 turbidity

recommended symbol: τ

The apparent absorbance of the incident radiation due to scattering.

Note

For small particles (see Definition 3.3.1), direct proportionality exists between turbidity and the Rayleigh ratio (see Definition 3.3.7).

3.3.10 particle scattering function, particle scattering factor

recommended symbols: $P(\theta)$, P_θ

The ratio of the intensity of radiation scattered at an angle of observation θ (see Definition 3.3.3) to the intensity of radiation scattered at an angle zero, i.e. $P(\theta) \equiv R(\theta)/R(0)$.

3.3.11 Zimm plot

A diagrammatic representation of data on scattering from large particles (see Definition 3.3.2), corresponding to the equation

$$\frac{Kc}{\Delta R(\theta)} = \frac{1}{\bar{M}_w P(\theta)} + 2A_2 c + \dots,$$

and used for the simultaneous evaluation of the mass-average molar mass (see Definition 2.7) \bar{M}_w , the second virial coefficient of the chemical potential (see Definition 3.1.7) A_2 and (usually) the z-average radius of gyration (see Definition 1.9) $\langle s^2 \rangle_z^{1/2}$; c is the mass concentration of the solute, $\Delta R(\theta)$ the excess Rayleigh ratio (see Definition 3.3.8), and $P(\theta)$ the particle

<i>Term</i>	<i>Definition</i>
	scattering function (see Definition 3.3.10) that comprises (usually) the z-average radius of gyration. K depends on the solute, the temperature and the type of radiation employed.
	<i>Note</i>
	Several modifications of the Zimm plot are in frequent use; the most common one uses the excess scattering (see Definition 3.3.8) instead of the excess Rayleigh ratio.
3.3.12 Guinier plot	A diagrammatic representation of data on scattering from large particles (see Definition 3.3.2), obtained at different angles but at the same concentration, constructed by plotting $\log \Delta R(\theta)$ or $\log P(\theta)$ versus $\sin^2(\theta/2)$ or q^2 , and (usually) used for the evaluation of the radius of gyration (see Definition 1.9). $\Delta R(\theta)$ is the excess Rayleigh ratio (see Definition 3.3.8), $P(\theta)$ the particle scattering function (see Definition 3.3.10), θ the scattering angle (see Definition 3.3.3) and q the length of the scattering vector (see Definition 3.3.5).
3.3.13 Kratky plot	A diagrammatic representation of scattering data on large particles, obtained at different angles but at the same concentration, constructed by plotting $\sin^2(\theta/2) \cdot \Delta R(\theta)$ versus $\sin(\theta/2)$, or $q^2 \Delta R(\theta)$ versus q , and used for the determination of molecular shape. For explanation of symbols, see Definition 3.3.12.
3.3.14 dissymmetry of scattering	recommended symbol: $z(\theta_1, \theta_2)$ The ratio of two Rayleigh ratios (see Definition 3.3.7) for different angles of observation (see Definition 3.3.3), i.e. $z(\theta_1, \theta_2) = R(\theta_1)/R(\theta_2)$, $\theta_1 < \theta_2$.
	<i>Note</i>
	The angles must be specified; in light scattering it is customary to let $\theta_2 = 180^\circ - \theta_1$ and, most frequently, $\theta_1 = 45^\circ$ and $\theta_2 = 135^\circ$.
3.3.15 depolarization of scattered light	The phenomenon, due primarily to the anisotropy of the polarizability of the scattering medium, resulting from the fact that the electric vectors of the incident and scattered beams are not coplanar and that, therefore, light scattered from a vertically (horizontally) polarized incident beam contains a horizontal (vertical) component.
3.3.16 turbidimetric titration	The process in which a precipitant is added incrementally to a highly dilute polymer solution and the intensity of light scattered by, or the turbidity (see Definition 3.3.9) due to, the finely dispersed particles of the polymer-rich phase (see Definition 3.4.3) is measured as a function of the amount of precipitant added.

DEFINITIONS

<i>Term</i>	<i>Definition</i>
3.3.17 isorefractive	An adjective describing components of a multicomponent system having zero refractive index increments with respect to each other.
3.3.18 Mie scattering	Scattering of light by particles with size larger than approximately one-half of the wavelength of incident light.

Note

For homogeneous spheres, this phenomenon is rigorously described by the theory developed by Mie.

3.4 Separation

<i>Term</i>	<i>Definition</i>
3.4.1 fractionation	A process by means of which macromolecular species differing in some characteristic (chemical composition, relative molecular mass (see Definition 1.1), branching, stereoregularity, etc) are separated from each other.
3.4.2 polymer-poor phase, dilute phase	That phase of a two-phase equilibrium system, consisting of a polymer and low-molecular-weight material, in which the polymer concentration is lower.
	<i>Note</i> The use of the name 'sol phase' is discouraged.
3.4.3 polymer-rich phase, concentrated phase	That phase of a two-phase equilibrium system, consisting of a polymer and low-molecular-weight material, in which the polymer concentration is higher.
	<i>Note</i> The use of the name 'gel phase' is discouraged.
3.4.4 precipitation fractionation	A process in which a polymeric material, consisting of macromolecules differing in some characteristic affecting their solubility, is separated from solution into fractions by successively decreasing the solution power of the solvent, resulting in the repeated formation of a two-phase system in which the less soluble components concentrate in the polymer-rich phase (see Definition 3.4.3).
3.4.5 extraction fractionation	A process in which a polymeric material, consisting of macromolecules differing in some characteristic affecting their solubility, is separated from a polymer-rich (see Definition 3.4.3) phase into fractions by successively increasing the solution power of the solvent, resulting in the repeated formation of a two phase system in which the more soluble components concentrate in the polymer-poor phase (see Definition 3.4.2).

<i>Term</i>	<i>Definition</i>
3.4.6 size-exclusion chromatography	recommended abbreviation: SEC A separation technique in which separation mainly according to the hydrodynamic volume (see Definition 3.2.3) of the molecules or particles takes place in a porous non-adsorbing material with pores of approximately the same size as the effective dimensions in solution of the molecules to be separated.
gel-permeation chromatography	recommended abbreviation: GPC The term 'gel-permeation chromatography' should only be used when the porous non-adsorbing material is a gel. The term 'size exclusion chromatography' is preferred.
3.4.7 molar-mass exclusion limit, molecular-weight exclusion limit	The maximum value of the molar mass (see Definition 1.2) or molecular weight (see Definition 1.1) of molecules or particles, in a specific polymer-solvent system, that can enter into the pores of the porous non-adsorbing material used in size-exclusion chromatography (see Definition 3.4.6).
	<i>Note</i> For particles with molar mass or molecular weight larger than the exclusion limit the separation effect of the size-exclusion chromatography vanishes.
3.4.8 elution volume	The volume of a solvent passed, since the injection of the sample, through a size-exclusion chromatography (see Definition 3.4.6) bed at the time at which a specified signal of the detector has been recorded.
3.4.9 retention volume	The elution volume (see Definition 3.4.8) at the maximum concentration of an elution peak.
3.4.10 universal calibration	A calibration of a size-exclusion chromatography (see Definition 3.4.6) set-up based on the finding that the retention volume (see Definition 3.4.9) of a molecular or particulate species is a single-valued function of an appropriate size parameter of this molecule or particle, irrespective of its chemical nature and structure.
	<i>Note</i> The product of the intrinsic viscosity (see Definition 3.2.21) and molar mass (see Definition 1.2), $[\eta]M$, has been widely used as the size parameter.
3.4.11 spreading function	The normalized signal produced, as a function of elution volume (see Definition 3.4.8), at the outlet of a size-exclusion chromatography (see Definition 3.4.6) set-up, by an instantaneous injection of a uniform sample.

DEFINITIONS

<i>Term</i>	<i>Definition</i>
3.4.12 plate number	recommended symbol: N A characteristic of the efficiency of a size-exclusion chromatography (see Definition 3.4.6) set-up in terms of band broadening, defined as $N = (V_R/\sigma_v)^2$, where V_R is the retention volume (see Definition 3.4.9) of an individual low-molecular-weight compound, and σ_v is the corresponding half width at 60.7% peak height of the elution peak.
3.4.13 plate height	recommended symbol: H The length of a part of a size-exclusion chromatography (see Definition 3.4.6) bed corresponding to one plate, i.e. the length of the bed, L , divided by its plate number (see Definition 3.4.12), N , or $H = L/N$.

4 ALPHABETICAL INDEX OF TERMS

<i>Term</i>	<i>Symbol</i>	<i>Definition number</i>
angle of observation	θ	3.3.3
apparent molar mass	M_{app}	2.11
apparent molecular weight	$M_{r,app}$	2.11
apparent relative molecular mass	$M_{r,app}$	2.11
Archibald's method		3.2.15
average degree of polymerization	\bar{X}	2.12
bead-rod model		3.2.4
bead-spring model		3.2.5
branching index	g	1.25
characteristic ratio	C_N, C_∞	1.13
chi parameter	χ	3.1.12
co-solvency		3.1.15
compositional heterogeneity		2.1
concentrated phase		3.4.3
constitutional heterogeneity		2.2
continuous distribution function		2.13
continuously curved chain		1.21
contour length		1.14
copolymer micelle		1.28
cross-over concentration	c^*	3.1.2
cross-section		3.3.7
cumulative distribution function		2.13
degree of polymerization	X	1.3
depolarization of scattered light		3.3.15
differential distribution function		2.13
dilute phase		3.4.2
dilute solution		3.1.1
discrete distribution function		2.13
distribution function		2.13
dissymmetry of scattering	$z(\theta_1, \theta_2)$	3.3.14
elution volume		3.4.8
end-to-end distance	r	1.11, 1.12
end-to-end vector	\mathbf{r}	1.10
equilibrium sedimentation (method)		3.2.13

INDIVIDUAL MACROMOLECULES

<i>Term</i>	<i>Symbol</i>	<i>Definition number</i>
equilibrium sedimentation in a density gradient		3.2.16
equivalent chain		1.17
equivalent sphere		1.4, 3.2.2
excess Rayleigh ratio		3.3.8
excess scattering		3.3.8
excluded volume of a macromolecule		3.1.9
excluded volume of a segment		3.1.8
expansion factor	$\alpha_r, \alpha_s, \alpha_c$	3.1.10
extraction fractionation		3.4.5
Flory constant	<i>F</i>	3.2.24
Flory distribution		2,17
Flory-Huggins theory		3.1.11
flow birefringence		3.2.9
fractionation		3.4.1
freely draining		3.2.6
freely jointed chain		1.16
freely rotating chain		1.19
frictional coefficient	<i>F</i>	3.2.1
gel-permeation chromatography		3.4.6
Guinier plot		3.3.12
Huggins coefficient	K_H	3.2.23
Huggins equation		3.2.22
hydrodynamic volume		3.2.3
hydrodynamically equivalent sphere		3.2.2
inherent viscosity	H_{inh}	3.2.20
integral distribution function		2.13
intrinsic viscosity	[<i>h</i>]	3.2.21
isopycnic		3.1.17
isorefractive		3.3.17
Kratky plot		3.3.13
Kuhn-Mark-Houwink-Sakurada equation		3.2.25
large particle		3.3.2
length of the scattering vector	<i>q</i>	3.3.5
limiting viscosity number	[<i>h</i>]	3.2.21
logarithmic normal distribution		2.20
logarithmic viscosity number	H_{ln}	3.2.20
long-chain branch		1.24
long-range interaction		1.6
long-range intramolecular interaction		1.6
Mark-Houwink equation		3.2.25
mass-distribution function		2.15
mass-average degree of polymerization	\bar{X}_w	2.12
mass-average molar mass	\bar{M}_w	2.7
mass-average relative molecular mass	$\bar{M}_{r,w}$	2.7
microgel		1,27
Mie scattering		3.3.18
molal refractive index increment		3.3.6
molar mass	<i>M</i>	1.2
molar-mass average	\bar{M}	2.5
molar-mass exclusion limit		3.4.7
molecular weight	M_r	1.1
molecular-weight average	\bar{M}_r	2.5
molecular-weight exclusion limit		3.4.7
monodisperse polymer		2.3

DEFINITIONS

<i>Term</i>	<i>Symbol</i>	<i>Definition number</i>
most probable distribution		2.17
network		1.26
non-draining		3.2.7
non-uniform polymer		2.4
number-average degree of polymerization		2.12
number-average molar mass	\bar{X}_n	2.6
number-average molecular weight	\bar{M}_n	2.6
number-average relative molecular mass	$\bar{M}_{r,n}$	2.6
number-distribution function		2.14
partially draining		3.2.8
particle scattering factor	$P(\mathbf{q}), P_q$	3.3.10
particle scattering function	$P(\mathbf{q}), P_q$	3.3.10
persistence length	a	1.22
perturbed dimensions		1.8
plate height	H	3.4.13
plate number	N	3.4.12
Poisson distribution		2.18
polydisperse polymer		2.4
polymer-poor phase		3.4.2
polymer-rich phase		3.4.3
polymer-solvent interaction		3.1.3
polymolecularity correction		2.21
Porod-Kratky chain		1.21
precipitation fractionation		3.4.4
preferential sorption		3.1.13
probability density function		2.13
quality of solvent		3.1.4
radius of gyration	$s, \langle s^2 \rangle^{1/2}$	1.9
random coil		1.15
random-walk chain		1.16
Rayleigh ratio	$R(\mathbf{q}), R_q$	3.3.7
reduced viscosity		3.2.19
refractive increment	$\mathcal{J}n/\mathcal{J}C$	3.3.6
refractive index increment	$\mathcal{J}n/\mathcal{J}C$	3.3.6
relative molecular mass	M	1.1
relative molecular-mass average	\bar{M}_r	2.5
relative viscosity	\mathbf{h}_r	3.2.17
relative viscosity increment	\mathbf{h}_i	3.2.18
retention volume		3.4.9
root-mean-square end-to-end distance	$\langle r^2 \rangle^{1/2}$	1.12
rotational diffusion		3.2.10
scattering angle	\mathbf{q}	3.3.3
scattering vector		3.3.4
Schulz-Flory distribution		2.17
Schulz-Zimm distribution		2.16
second virial coefficient	A_2	3.1.7
sedimentation coefficient	s	3.2.11
sedimentation equilibrium		3.2.12
sedimentation velocity method		3.2.14
selective sorption		3.1.13
selective solvent		3.1.14
short-chain branch		1.23
short-range interaction		1.5
short-range intramolecular interaction		1.5

<i>Term</i>	<i>Symbol</i>	<i>Definition number</i>
size-exclusion chromatography		3.4.6
small particle		3.3.1
solubility parameter	d	3.1.16
specific refractive index increment		3.3.6
spreading function		3.4.11
statistical coil		1.15
statistical segment		1.18
Staudinger index	[h]	3.2.21
steric factor	s	1.20
streaming birefringence		3.2.9
thermodynamic quality of solvent		3.1.4
thermodynamically equivalent sphere		1.4
theta solvent		3.1.5
theta state		3.1.5
theta temperature		3.1.6
Tung distribution		2.19
turbidimetric titration		3.3.16
turbidity	t	3.3.9
uniform polymer		2.3
universal calibration		3.4.10
unperturbed dimensions		1.7
virial coefficients		3.1.7
virial coefficients of the chemical potential	A_1, A_2, \dots	3.1.7
viscosity function	F	3.2.24
viscosity number		3.2.19
viscosity ratio	h_r	3.2.17
viscosity-average degree of polymerization	\bar{X}_v	2.12
viscosity-average molar mass	\bar{M}_v	2.10
viscosity-average molecular weight	$\bar{M}_{r,v}$	2.10
viscosity-average relative molecular mass	$\bar{M}_{r,v}$	2.10
weight-average degree of polymerization	\bar{X}_w	2.12
weight-average molecular weight	$\bar{M}_{r,w}$	2.7
weight-distribution function		2.15
worm-like chain		1.21
Zimm plot		3.3.11
z-average degree of polymerization	\bar{X}_z	2.12
z-average molar mass	\bar{M}_z	2.8
z-average molecular weight	$\bar{M}_{r,z}$	2.8
z-average relative molecular mass	$\bar{M}_{r,z}$	2.8
(z+1)-average degree of polymerization	\bar{X}_{z+1}	2.12
(z+ 1)-average molar mass	\bar{M}_{z+1}	2.9
(z+1)-average molecular weight	$\bar{M}_{r,z+1}$	2.9
(z+1)-average relative molecular mass	$\bar{M}_{r,z+1}$	2.9

REFERENCES

- 1 IUPAC. Basic definitions of terms relating to polymers 1974. *Pure Appl. Chem.* **40**, 477-491 (1974). Reprinted as Chapter 1, this volume.
- 2 IUPAC. Stereochemical definitions and notations relating to polymers (recommendations 1980). *Pure Appl. Chem.* **53**, 733-752 (1981). Reprinted as Chapter 2, this volume.
- 3 IUPAC. Note on the terminology for molar masses in polymer science. *Makromol. Chem.* **185**, Appendix to No. 1 (1984); *J. Polym. Sci., Polym. Lett. Ed.* **22**, 57 (1984); and in other journals.