stoichiometric mean molal activity coefficient (practical activity coefficient) (in electrochemistry)

The practical activity coefficient of electrolyte B is given by

$$\gamma_{\pm} = a_{\pm}/(\nu_{+}^{\nu_{+}}\, \nu_{-}^{\nu_{-}})^{1/2}(m_B/m^{+})$$

where $a_{\pm}$ is the mean activity of B in solution, m_B is the *molality* of B, m^{+} = 1 mol kg^{-1}, ν_{+} is the number of cations and ν_{-} the number of anions in the chosen group B which is taken as the electrolyte.

$$\nu = \nu_{+} + \nu_{-}$$

1974, *37*, 510