

standard electromotive force

Quantity defined by $E^{\circ} = -\Delta_{\text{r}}G^{\circ}/nF = (RT/nF) \ln K^{\circ}$, where $\Delta_{\text{r}}G^{\circ}$ is the standard Gibbs energy of the cell reaction in the direction in which reduction occurs at the right-hand electrode in the diagram representing the cell ('reduction at right'), K° is the standard equilibrium constant for this reaction, n its charge number, F the Faraday constant, R the gas constant and T the thermodynamic temperature.

G.B. 58