

Pauli exclusion principle

A rule, complementary to the *aufbau* principle, of building up the electronic configuration of atoms and molecules: a maximum of two electrons can occupy an orbital and then only providing that the spins of the electrons are paired, i.e. opposed. The principle demands that the wavefunction for a many-electron system must be asymmetric with respect to the permutation of the space-spin coordinates for every pair of electrons.

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